

## Ionization Energy and Electronegativity Table of the Elements

314 <b>H*</b> 2.1																	566 <b>He*</b>	
													<b>1st Ionization Energy (kcal/mol)</b> →					
													<b>Electronegativity (fluorine = 4.0)</b> →					
124 <b>Li</b> 1.0	215 <b>Be</b> 1.5											191 <b>B</b> 2.0	260 <b>C</b> 2.5	335 <b>N*</b> 3.0	312 <b>O*</b> 3.5	402 <b>F*</b> 4.0	498 <b>Ne*</b>	
119 <b>Na</b> 0.9	176 <b>Mg</b> 2											138 <b>Al</b> 1.5	188 <b>Si</b> 1.8	254 <b>P</b> 2.1	239 <b>S</b> 2.5	300 <b>Cl*</b> 3.0	363 <b>Ar*</b>	
100 <b>K</b> 0.8	141 <b>Ca</b> 1.0	151 <b>Sc</b> 1.3	158 <b>Ti</b> 1.5	155 <b>V</b> 1.6	156 <b>Cr</b> 1.6	171 <b>Mn</b> 1.5	182 <b>Fe</b> 1.8	181 <b>Co</b> 1.8	176 <b>Ni</b> 1.8	178 <b>Cu</b> 1.9	216 <b>Zn</b> 1.6	138 <b>Ga</b> 1.6	187 <b>Ge</b> 1.8	242 <b>As</b> 2.0	225 <b>Se</b> 2.4	273 <b>Br**</b> 2.8	323 <b>Kr*</b>	
96 <b>Rb</b> 0.8	131 <b>Sr</b> 1.0 <small>Strontium</small>	152 <b>Y</b> 1.2	160 <b>Zr</b> 1.4	156 <b>Nb</b> 1.6	166 <b>Mo</b> 1.8	172 <b>Tc</b> 1.9	173 <b>Ru</b> 2.2	178 <b>Rh</b> 2.2	192 <b>Pd</b> 2.2	174 <b>Ag</b> 1.9	207 <b>Cd</b> 1.7	133 <b>In</b> 1.7	169 <b>Sn</b> 1.8	199 <b>Sb</b> 1.9	208 <b>Te</b> 2.1	241 <b>I</b> 2.5	280 <b>Xe*</b>	
90 <b>Cs</b> 0.7	120 <b>Ba</b> 0.9	129 <b>La</b> 1.1	127 <b>Hf</b> 1.3	140 <b>Ta</b> 1.5	184 <b>W</b> 1.7	181 <b>Re</b> 1.9	201 <b>Os</b> 2.2	212 <b>Ir</b> 2.2	208 <b>Pt</b> 2.2	212 <b>Au</b> 2.4	241 <b>Hg**</b> 1.9	141 <b>Tl</b> 1.8	171 <b>Pb</b> 1.8	184 <b>Bi</b> 1.9	196 <b>Po</b> 2.0	248 <b>At</b> 2.2	248 <b>Rn*</b>	
Fr 0.7	122 <b>Ra</b> 0.9	162 <b>Ac</b> 1.1																

159 <b>Ce</b> 1.1	133 <b>Pr</b> 1.1	145 <b>Nd</b> 1.1	Pm 1.1	129 <b>Sm</b> 1.1	131 <b>Eu</b> 1.1	142 <b>Gd</b> 1.1	155 <b>Tb</b> 1.1	157 <b>Dy</b> 1.1	Ho 1.1	Er 1.1	Tm 1.1	143 <b>Yb</b> 1.1	115 <b>Lu</b> 1.1
Th 1.3	Pa 1.5	92 <b>U</b> 1.7	Np 1.3	Pu 1.3	Am 1.3	Cm 1.3	Bk 1.3	Cf 1.3	Es 1.3	Fm 1.3	Md 1.3	No 1.3	Lr

**Key:**  
**Solid**  
 \* **Gas**  
 \*\* **Liquid**  
 Not Found in Nature